

2. For the following substances:

- 1) in the first column, class the compound as ionic (i), molecular (m), acid (a) or hydrated ionic (h).
- 2) give the correct chemical *formula* in the space provided; some formulas correspond to traditional names.

i, m, a or h	Name	Formula
1 A	chloric acid	$\text{HClO}_3 (\text{aq})$
2 I	strontium sulfide	SrS
3 I	manganese(II) sulfate	MnSO_4
4 I	manganese(IV) sulfate	$\text{Mn}(\text{SO}_4)_2$
5 m	ethanol	$\text{C}_2\text{H}_5\text{OH}$
6 H	cadmium thiosulfate dihydrate	$\text{CdS}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$
7 I	aluminum permanganate	$\text{Al}(\text{MnO}_4)_3$
8 A	boric acid	$\text{H}_3\text{BO}_3 (\text{aq})$
9 I	magnesium benzoate	$\text{Mg}(\text{C}_6\text{H}_5\text{COO})_2$
10 I	cesium nitride	Cs_3N
11 I	chromium(III) oxide	Cr_2O_3
12 m	trinitrogen tetrachloride	N_3Cl_4
13 m	ammonia	NH_3
14 I	ammonium hydroxide	NH_4OH
15 I	lithium borate	Li_3BO_3
16 m	bromine	Br_2
17 m	dinitrogen tetrahydride	N_2H_4
18 I	sodium acetate	NaCH_3COO
19 I	calcium cyanide	$\text{Ca}(\text{CN})_2$
20 m	silicon dioxide	SiO_2
21 m	hydrogen peroxide	H_2O_2
22 A	acetic acid	$\text{CH}_3\text{COOH} (\text{aq})$
23 I	sodium nitrite	NaNO_2
24 I	sodium nitrate	NaNO_3
25 I	calcium nitrite	$\text{Ca}(\text{NO}_2)_2$
26 I	calcium nitrate	$\text{Ca}(\text{NO}_3)_2$
27 I	zinc hydrogen phosphate	ZnHPO_4
28 I	zinc phosphate	$\text{Zn}_3(\text{PO}_4)_2$
29 I	zinc phosphide	Zn_3P_2
30 m	ozone	O_3
31 m	sucrose	$\text{C}_{12}\text{H}_{22}\text{O}_{11}$
32 m	trichlorine dioxide	Cl_3O_2
33 m	carbon disulfide	CS_2
34 I	aluminum chloride	AlCl_3
35 I	gallium hydroxide	$\text{Ga}(\text{OH})_3$