

8. Complete the following table:

Formula	Ionic or Molecular (I or M)	Name
N_3P_5		
$CaCl_2$		
$MgCO_3 \cdot 2H_2O$		
$NaNO_3$		
SO_2		
P_3O_3		

9. Write the formula for the following molecular compounds:

- TetraSulfur Nonachloride
- Nitrogen TriBromide
- Carbon Monoxide
- Octaphosphorous Decahydride

10. Write the formulas for the following ionic compounds

- Magnesium Fluoride
- Potassium Oxide
- Calcium Oxide
- Mercury(I)Iodide
- Mercury (II) Iodide
- Sodium Chlorate

11. More practice with naming ionic compounds

- Iron(II) Nitrate
- Copper(I) Phosphate
- Beryllium Sulfate
- Sodium Acetate
- Sodium Sulfide Dihydrate
- Magnesium Chloride Monohydrate
- Radium Nitrite Heptahydrate
- Calcium Chlorate

12. Make a chart comparing the similarities of Acids and Bases

13. Name the following acids:

- $HBr(aq)$
- $HCN(aq)$
- $HCH_3COO(aq)$
- $HClO_2(aq)$
- $HClO_3(aq)$
- $HClO_4(aq)$
- $HCl(aq)$